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Thermodynamic behavior of $\text{FeCl}_3\text{-H}_2\text{O}$ and $\text{HCl-FeCl}_3\text{-H}_2\text{O}$ systems - a Pitzer model at 25 °C

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Abstract

Most of the speciation models of aqueous electrolyte systems are based on the “ion pairing and complexing approach”. They are only suitable for aqueous solutions with ionic strengths lower than about 1 eq.kg_w^{-1} . For highly saline solutions, an approach based on “specific-ion interactions” (the so-called “Pitzer approach”) provides much more accurate results.

Various thermodynamic databases supporting the “Pitzer approach” exist, but they are often inconsistent and poorly documented. The number of chemical elements is limited because of the difficulties inherent to producing relevant experimental data for wide ranges of temperature, ionic strength, pH, and for low solubility metals and hydroxides, and also because of the difficulty in evaluation of consistent and accurate values for interaction parameters.

This study presents a new set of parameters using the standard Pitzer formalism that is able to model the behavior of Fe-rich solutions and solid-liquid equilibria in binary $\text{FeCl}_3\text{-H}_2\text{O}$ and mixed $\text{HCl-FeCl}_3\text{-H}_2\text{O}$ systems at 25 °C. The new binary and mixing interaction parameters are adjusted in order to fit various independent experimental activity (water activity, osmotic coefficients) and ferric chloride salts solubility data. This work allows updating and improving thermodynamic databases for Fe(III) systems in order to simulate the complex geochemical behaviour of solutions and gas-brine-rock interactions in various applications in the geosciences or industrial processes.

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Keywords: acidic saline solutions; interaction parameters; Pitzer; ferric chloride salts.

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1. Introduction

Knowledge of the thermodynamic behavior of saline multicomponent aqueous systems is important in industrial processes as well as in geosciences. Increasing use of such non-conventional water resources needs development of robust geochemical models for describing and predicting chemical reactions within these particular media (dissolution/precipitation of salts, evaporation...). The Pitzer approach, usually quoted to describe such complex systems, is suitable to deal with major elements in brines. But, some systems are poorly documented and the Pitzer databases have to be continuously improved to integrate new chemical species and interaction parameters.

This study focuses on very acidic systems and on the efforts made to define the interactions of Fe(III) species within the H-Fe-Cl-H₂O system at 25°C.

2. Bibliography and critical selection of experimental data

The H-Fe(III)-Cl-H₂O system was investigated by numerous authors both the experimental and numerical aspects. For the experimental part, water activity data are presented for the binary FeCl₃-H₂O system [1-2], as well as in mixed HCl-FeCl₃-H₂O solutions [3-6]. In these works, water activities were measured in solutions up to concentrations of 7 and 4 m for HCl and FeCl₃, respectively.

Thermodynamic models for the Fe(III) chloride binary system, based on the Pitzer formalism, were developed by various authors, each of them establishing a different set of parameters [7-10]. It should be noted that to accurately describe the activity properties of FeCl₃-H₂O system at high molality, the models used the extended Pitzer approach with four binary parameters (β^0 , β^1 , β^2 , c^*) [8-9]. The model of Marion [10] extends to low temperature the 25°C model of Christov [9]. All these models [8-10] give a very good agreement with both osmotic coefficients, water activity, and FeCl₃:6H₂O_(cr) solubility data in pure water at 25°C { $m(\text{sat}) = 6.1 \text{ mol.kg}_w^{-1}$ at 25°C}. A model for the mixed HCl-FeCl₃-H₂O system was also developed by [8], which can describe accurately the solubility of FeCl₃:6H₂O_(cr) up to 5 M of HCl.

However, in mixed hydrochloric solutions the solubility of the ferric chloride salt increases sharply with increasing HCl concentration [11-12]. According to the experimental solubility data, in addition to hexahydrate (FeCl₃:6H₂O_(cr)), low crystal hydrates (FeCl₃:3.5H₂O, FeCl₃:2.5H₂O, FeCl₃:2H₂O) and double salt (FeCl₃.HCl.2H₂O) precipitate from saturated mixed HCl-FeCl₃-H₂O solutions. These low crystal hydrates have extremely high solubilities, greater than 13 mol.kg_w⁻¹ in solutions with relatively low acidity (between 1 and 5 mol.kg_w⁻¹ of HCl). Such solubility corresponds to an ionic strength greater than 75 m, if the salt is assumed totally dissociated. This extremely high solubility of ferric chloride salts in mixed systems leads to limitations of the solid-liquid equilibrium models for the FeCl₃-H₂O system available from the literature [7-9]. Current binary and ternary parameters are not sufficiently “robust” and the models are not able to reproduce accurately such salt solubilities.

Within this study we evaluate new sets of pure electrolyte parameters of Fe(III)-Cl interactions, which are valid up to a very high molality supersaturated (with respect to the stable FeCl₃:6H₂O_(cr) solid) FeCl₃-H₂O solutions. The new FeCl₃-H₂O binary parameters are then used to develop a high molality solid-liquid equilibrium model for mixed HCl-FeCl₃-H₂O system. Two ferric chloride salts (FeCl₃:3.5H₂O and FeCl₃:2.5H₂O) are added to the database and their respective equilibrium constants at 25°C are also determined. The defined parameters are developed for the PHREEQC geochemical code [13], and according to the specific formalism implemented in the program.

3. Results of numerical modeling

3.1. Binary parameters for $\text{FeCl}_3\text{-H}_2\text{O}$ system

The new set of parameters for the binary $\text{FeCl}_3\text{-H}_2\text{O}$ system is established using the experimental data given in [1-2]. The estimation procedure is described in [14]. As done in [8-9], a 4-parameter approach is applied, which allows fitting the osmotic coefficient data without considering the existence of a complex species in solution (as FeCl^{2+} , FeCl_2^+ , $\text{FeCl}_3(\text{aq})$, as suggested in [4-5]).

The resulting parameterization gives good agreement with experimental osmotic coefficients data ($\sigma = 0.0106$). The new model permits inclusion of three ferric chloride minerals (Table 1), which precipitate in the $\text{FeCl}_3\text{-H}_2\text{O}$ system at extremely high molality (> 16 m). The calculated equilibrium constants give $m(\text{sat})$ predictions which are in excellent agreement with experimental solubility data (Table 1).

Table 1. Experimental and predicted solubilities of ferric chloride salts in binary $\text{FeCl}_3\text{-H}_2\text{O}$ system at 25°C

Solid phases	Experimental $m(\text{sat})$ (mol/kg _w) in [12]	Calculated $m(\text{sat})$ (mol/kg _w) (this study)
$\text{FeCl}_3 \cdot 6 \text{H}_2\text{O}$	6.05	6.05
$\text{FeCl}_3 \cdot 3.5 \text{H}_2\text{O}$	13.04	13.26
$\text{FeCl}_3 \cdot 2.5 \text{H}_2\text{O}$	16.10	16.12

3.2. Ternary parameters for $\text{HCl-FeCl}_3\text{-H}_2\text{O}$ system

The ternary interaction parameters ($\Theta_{\text{H-Fe}}$ and $\Psi_{\text{H-Fe-Cl}}$) are evaluated using the experimental $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}(\text{cr})$ solubility data in $\text{HCl-FeCl}_3\text{-H}_2\text{O}$ system [3]. The $\text{HCl-H}_2\text{O}$ parameterization (binary parameters) is issued from [15] and it is valid up to ≈ 20 m at 25°C [16]. The new mixed system parameterization provides a very good agreement between model predictions and:

- *the water equi-activity measurements* (Fig. 1a). The established parameterization gives very good agreement both in high and low concentration solutions;
- *the composition of invariant (eutonic) solutions*. The new set of parameters can reproduce the measured composition of eutonic solutions with a maximum difference in molality of about 5% (Table 2);
- *the measured pH of acid solutions* (Fig. 1b), even in the range of negative values of pH.

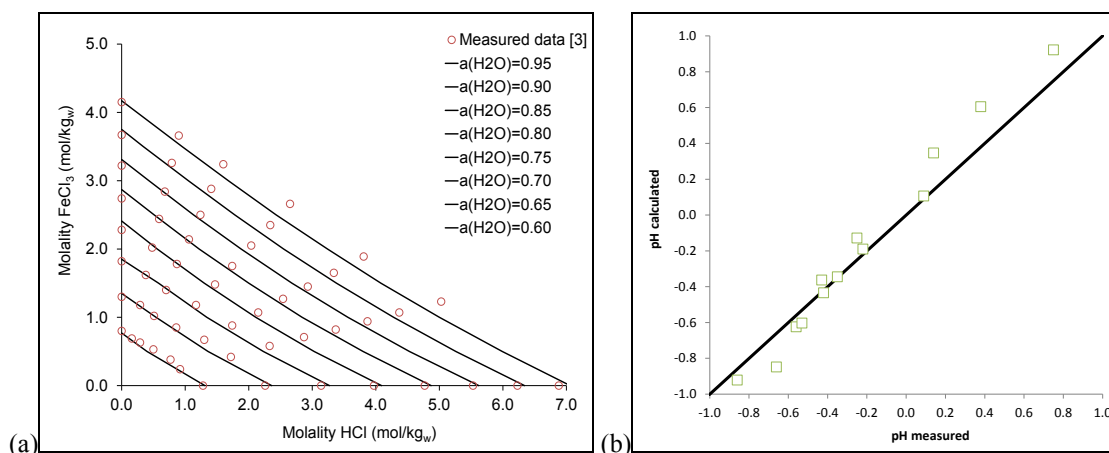


Fig. 1 In HCl-FeCl₃-H₂O solutions at 25°C: (a) A comparison between experimental (red circles, [3]) and calculated (solid lines) water activities; b) A comparison between calculated (solid line) and measured (open squares, [17]) pH values.

Table 2. Experimental and predicted composition of invariant points in HCl-FeCl₃-H₂O systems at 25°C

Solid phases	Experimental data [11-12]		Calculated results (This study)		
	HCl (mol/kgw)	FeCl ₃ (mol/kgw)	HCl (mol/kgw)	FeCl ₃ (mol/kgw)	Ionic strength
FeCl ₃ ·6H ₂ O + FeCl ₃ ·3.5H ₂ O	1.29	13.18	1.29	12.47	75.8
FeCl ₃ ·3.5H ₂ O + FeCl ₃ ·2.5H ₂ O	4.16	16.51	4.16	17.47	109.0

4. Concluding remarks

The set of parameters established in this work is suitable to reproduce water characteristics of ferric chloride solutions both at low and high ionic strengths. This new set of parameters can represent the composition of highly saline solutions up to the eutonic points of hydrated ferric chloride salts in the mixed HCl-FeCl₃-H₂O system. All binary and mixed solution interaction parameters and equilibrium constants of Fe(III) minerals are coherent, and therefore can be used to determine the activities of dissolved species up to the salt solubility concentrations. The models have been developed in order to update the PHREEQC code. The presented parameterization can be used without any additional adjustments to determine the solution behaviour of more complex Fe(III) systems, as well as to develop a solubility model for a very important geochemical Fe(III) hydroxides minerals in high salinity solutions.

References

- [1] Rumyantsev AV, Hagemann S, Moog HC. Isopiestic investigation of the systems Fe₂(SO₄)₃-H₂SO₄-H₂O, FeCl₃-H₂O, and Fe(III)-(Na, K, Mg, Ca)Cln-H₂O at 298.15 K. *Zeitschrift für Physikalische Chemie* 2004; **218**: 1089-127.
- [2] Mikulin G. *Voprosy Fizicheskoi Khimii Electrolytov*. Izd. Khimiya, 1968.
- [3] Kangro W, Groeneveld A. Konzentrierte wäßrige Lösungen, I. *Z Phys Chem (Frankfurt am Main)* 1962; **32**: 110-26.
- [4] Majima H, Awakura Y. Water and solute activities of H₂SO₄-Fe₂(SO₄)₃-H₂O and HCl-FeCl₃-H₂O solution systems: Part I. Activities of water. *Metallurgical Transactions B* 1985; **16B**: 433-9.
- [5] Majima H, Awakura Y. Water and solute activities of H₂SO₄-Fe₂(SO₄)₃-H₂O and HCl-FeCl₃-H₂O solution systems: Part II. Activities of solutes. *Metallurgical Transactions B* 1986; **17B**: 621-7.
- [6] Awakura Y, Kawasaki Y, Uno A, Sato K, Majima H. Activities of water and HCl in aqueous solution systems of HCl-MCl_n including CuCl₂, NiCl₂ and FeCl₃. *Hydrometallurgy* 1987; **19**: 137-57.
- [7] Kim HT, Frederick WJ. Evaluation of ion interaction parameters of aqueous electrolytes at 25°C. 1. Single salt parameters. *J Chem Eng Data* 1988; **33**:177-84.
- [8] Tanaka M, Tamagawa T, Hamada Y. Estimation of activities in the aqueous solution systems of HCl-CuCl₂ and HCl-FeCl₃ using the Pitzer method. *Materials Transactions, JIM* 1992; **33(4)**: 391-9.
- [9] Christov C. Pitzer ion-interaction parameters for Fe(II) and Fe(III) in the quinary {Na + K + Mg + Cl + SO₄ + H₂O} system at T = 298.15 K. *J Chem Thermodynamics* 2004; **36**: 223-235.
- [10] Marion GM, Kargel JS, Catling DC. Modeling ferrous-ferric iron chemistry with application to martian surface geochemistry. *Geochim Cosmochim Acta* 2008; **72**: 242-66.
- [11] Gmelin's Handbuch der Anorganischen Chemie, Eisen, Al [B]. Chemie, Berlin, 1932.
- [12] Linke W. *Solubilities Inorganic and Metal-Organic Compounds*, Vols 1 and 2. 4th ed. Washington: American Chemical Society; 1965.

- [13] Parkhurst DL, Appelo CAJ. *User's guide to PHREEQC (version 2) - A computer program for speciation, batch-reaction, one-dimensional transport, and inverse geochemical calculations*. U.S. Geological Survey Water-Resources Investigations Report 1999; 99-4259.
- [14] André L, Lassin A, Azaroual M. A methodology to estimate pitzer's interaction parameters. *Geochim Cosmochim Acta* 2009; **73(13), Suppl. 1**: A41.
- [15] Christov C, Moller N. Chemical equilibrium model of solution behavior and solubility in the H-Na-K-OH-Cl-HSO₄-SO₄-H₂O system to high concentration and temperature. *Geochim Cosmochim Acta* 2004; **68**: 1309-31.
- [16] Christov C. Chemical equilibrium model of solution behavior and solubility in the MgCl₂-H₂O, and HCl-MgCl₂-H₂O systems to high concentration from 0°C to 100°C. *J Chem Eng Data* 2009; **54**: 2599-608.
- [17] See MS. Use of the Bromley equation for the analysis of ionic equilibria in mixed ferric and ferrous chloride solutions at 25°C. *Metallurgical and Materials Transactions B* 2006; **37B**: 173-9.